Unit 4 Major Quiz Problem-Based Review

These questions are purely calculation/problem based. This sheet is not sufficient review on its own. Be sure to \*also\* study the periodic table and relevant history, and check out the Kahoots for more practice.

1. Analyze each isotope and fill in the chart.

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| Symbol-  Mass # | % Abundance | Mass of Atom | # protons | # electrons | # neutrons | Isotopic Notation |
| K-39 | 93.26% | 38.96 amu |  |  |  |  |
| K-40 | 0.01% | 39.96 amu |  |  |  |  |
|  | 6.73% | 40.96 amu | 19 |  | 22 |  |

2. Calculate the average atomic mass of potassium using the information in the previous

chart.

3. Europium has two stable isotopes, 151Eu (atomic mass 150.9198 u) and 153Eu (atomic mass

152.9212 u). If the average atomic mass of europium is 151.97 u, what are the percent

abundances of each isotope?

4. To eject one electron from the surface of potassium metal, a potassium atom must absorb

2.3 eV in the form of a quantum.

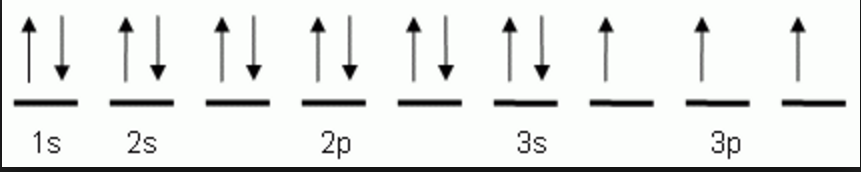
1. Calculate the wavelength, in nm, of this quantum.
2. Calculate the frequency of this quantum.
3. How much energy, in Joules, is required to eject a mole of electrons from the surface of potassium metal?

5. A photon of yellow light has a wavelength of about 570 nm.

1. Calculate the frequency of yellow light?
2. Does yellow light have sufficient energy to excite a hydrogen electron from its ground state of n = 2 to n = 5, which requires the absorption of 4.58 x 10-19 J?

6. Are these sets of quantum numbers allowable?

1. 3, 1, -2, +½ b) 4, 2, -2, -½ c) 3, 2, 0, -½ d) 5, 6, -4, +1/2



7. Write an allowable set of quantum numbers

for the indicated electron.

For the following questions, note that exceptions to the Aufbau order may exist.

8. Write the unabbreviated electron configuration for each element, determine their number of

valence electrons, and its valence energy level.

1. F
2. V
3. Sn

9. Write the noble gas core electron configuration for:

1. Nd
2. Cd

10. Identify the element:

1. [Ar] 4s13d10
2. [Xe] 6s2 5d7

11. Are these atoms and/or ions isoelectronic? To prove this, write the configurations first.

1. N3- and Sc+3
2. Kr and Se2-

12. Draw the noble gas core orbital notation for:

1. Nb
2. As
3. Np
4. Ag

13. Draw the noble gas core orbital notation - valence orbitals only - for:

1. N b) Sb